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CHEM201A Chapter 6 Lecture Outline Gottlieb 15 Example: Determine the mass of hydrogen gas, in g, that must be combusted to produce 1510.2 kJ of heat.  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$   $\Delta H_{\text{rxn}} = -483.6 \text{ kJ}$  Example: A 32.3-mL sample of 0.996M NaOH is mixed with 19.1 mL of 0.905M H<sub>2</sub>SO<sub>4</sub>.

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